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## Elements definition chemistry class 9

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Features: • It may be heterogeneous. • The properties of the ingredients are preserved. • No new compound is formed. • Elements can be separated by simple physical processes. • It does not have a fixed melting and boiling point. Type 1 of the method of separating the mixture. Two immiscible liquids. Example of oil + water Using the separating funnel 2. Ammonium chloride + sand Sublimation 3. Dyes in black ink. (One solvent different ingredients) Chromatography 4. Two different liquids. (A stone + water) Distillation 5. Solid particles insoluble (solvent) liquid. (Milk + cream) Centrifugation Physical and chemical change Physical change Chemical change (i) No new substance is formed. A new substance is forming. (ii) The properties of the ingredients/substances are preserved. Characteristics of changes in ingredients/substances. (iii) The change does not involve loss of heat or profit. This reaction may include heat loss or profit. (iv) This amendment is normally reversible. This change is usually irreversible. • Solution It is a homogeneous mixture of two or more substances. Solvent A substance dissolved in a solvent, e. g. Liquid part of the solution to be dissolved, e.g. Soluble can be solid, liquid or gas. The solvent may be liquid, solid or gas. Suspension Solvent (i) The size of soluble particles is visible on open eyes The size of soluble particles is not visible on open eyes. (ii) Indicates the effect of the tyndall Indicates the effect of the tyndall (iii) Translucent translucent (iv) Soluble particles descending Colloidal particles do not descend • Different types of colloids Scattered phase dispersion Instrument type Example of liquid gas aerosol mist, clouds, mist Solid gas Aerosol Smoke, car exhaust Gas Liquid Foam Shaving cream Liquid Liquid Emulsion Milk, Face Cream Solid Liquid Liquid Solution Magnesia Milk, Mud Gas Solid Goam Fungus, Puss fluid Solid Gel Jelly, Cheese, Butter Solid Solid Sol Colored Gemstone, Milky Glass What Is Element? a substance whose atoms all have the same number of protons; another way of saying this is that all atoms of a particular element have the same serial number. Elements are chemically simplest substances, so they cannot be separated by chemical reactions. Elements can only be changed to other elements using kernel methods. Although the atoms of the element must have the same number of protons, they can have a different number of neutrons and therefore different masses. When atoms of the same element have a different number of neutrons, they are called isotopes. How the elements were correctly defined In 1913, chemistry and physics were topsy-turvy. Some great hitters - including Dmitri Mendeleev - spoke seriously of elements lighter than hydrogen and elements between hydrogen and helium. The visualization of the atom was free for all, and Mendeleev's justification for the inter-atomic lock based on the atomic weights of the initials was falling apart at the seams. This is the story of how Henry Moseley brought light into the darkness. With only one proton, hydrogen is the simplest, lightest element, followed by helium with two protons. The oxygen atom has eight protons. At 75%, hydrogen is the most abundant element in the universe, followed by helium at 23% and oxygen at 1%. All the other elements make up the remaining 1 percent. Oxygen in the earth's crust (47%) is the most abundant element followed by silicon (28%) and aluminium (8%). Element names and numbers All elements have been named. Some of these names are familiar to us, such as nitrogen and sodium, and some are less familiar, such as dysprosium and roentgenium. We can also name elements by means of their serial numbers. For example, element 1 is hydrogen, element 2 is helium, element 3 lithium, element 8 is oxygen, etc. How many elements are there? There are currently 118 approved elements. The period table is used to display all elements in an organized manner. Elements Ancient and modern Some elements have been known for thousands of years, we don't know who found them. These are: antimony, arsenic, coal, copper, iron, gold, lead, mercury, silver, sulfur and tin. All other elements have been discovered since 1669: this year, Hennig Brand became the first known person to discover a new element - phosphorus. Combining elements An element can merge with one or more other elements, forming compounds of which there are millions. For example, one of the most well-known compounds is water chemically written under the name H<sub>2</sub>O, which means that the water is made of two hydrogen atoms chemically glued to one oxygen. QuizNext learning material ICSE 9 Chemistry Let QuizNext's AI helps with accurate versioning. Take your daily quizzes and stay on top of the times! Answer: The element is a single type of atom, while the compound consists of two or more atomic types. Elements cannot be further divided into smaller units without using large quantities ... chloride (NaCl), a compound consisting of sodium (Na) and chlorine(c) Explanation: mark me as the most cerebral PLZ PLZ

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